

# Science

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(Chapter - 1) (Chemical Reactions and Equations) (Practice Test 2)

(Class X)

Time: 1 Hour

Maximum Marks: 25

## General Instructions:

- This question paper contains four sections – A, B, C, and D. Each part is compulsory.
- Section – A has 4 MCQ of one mark each.
- Section – B has 5 questions of two marks each.
- Section – C has 2 questions of three marks each.
- Section – D has 1 questions of five marks.

## Section - A

1. When Magnesium ribbon burns....Choose a more correct answer.
  - a) white powder is form
  - b) reaction was completed in the absence of oxygen
  - c) formation of magnesium oxide
  - d) ribbon burns with a dazzling white flame and changes into white powder
2. Action of dilute sulphuric acid on zinc.
  - a) formation of sulphuric zinc
  - b) zinc get more diluted
  - c) carbon dioxide formed
  - d) formation of hydrogen gas
3. Chemical change cannot be determined by.....
  - a) change in state
  - b) change in color
  - c) evolution of gas
  - d) bubbles come out
4. Choose the following observation which shows the chemical change.
  - a) Change in smell
  - b) Vapors comes out
  - c) Reaction color change
  - d) Reactant state change before taking part in reaction

## Section - B

5. Name the reducing agent in the following reaction:



State which is more reactive, *Mn* or *Al* and Why?

6. "We need to balance a skeleton chemical equation". Give reason to justify the statement.
7. Define oxidation and reduction.
8. Write a balanced equation for the chemical reaction that can be characterized as precipitation reaction.
9. Hydrogen being a highly inflammable gas and oxygen being a supporter of combustion, yet water, a compound made up of hydrogen and oxygen is used to extinguish fire. Why?

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Section - C

10. In the electrolysis of water:
- Name the gas collected at the cathode and anode respectively.
  - Why is the volume of gas collected at one electrode double than that at the other? Name this gas.
  - How will you test this gas?
11. State the type of chemical reactions and chemical equations that take place in the following:
- Magnesium wire is burnt in air.
  - Electric current is passed through water.
  - Ammonia and hydrogen chloride gasses are mixed.

Section - D

12. Identify the type of chemical reaction in the following statement and define each of them:
- Digestion of food in our body.
  - Rusting of iron.
  - Heating of manganese dioxide with aluminum powder
  - Blue color of copper sulfate solution disappears when iron filings are added to it.
  - Dilute hydrochloric acid is added to the sodium hydroxide solution to form sodium chloride and water.



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