

# Science

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## (Chapter - 1) (Chemical Reactions and Equations) (Practice Test 3) (Class X)

Time: 1 Hour

Maximum Marks: 25

### General Instructions:

- This question paper contains four sections – A, B, C, and D. Each part is compulsory.
- Section – A has 4 MCQ of one mark each.
- Section – B has 5 questions of two marks each.
- Section – C has 2 questions of three marks each.
- Section – D has 1 question of five marks.

### Section - A

1. Dilute sulphuric acid + ..... formation of hydrogen gas

- a) Zinc oxide
- b) Zinc granule
- c) Zinc chloride
- d) Zinc ribbon

2. Chemical change is determined by...

- a) Checking temperature with hand or thermometer
- b) Observing reactants
- c) Bubbles comes out
- d) Checking smell

3. .... + ..... Magnesium oxide

- a) Magnesium, hydrogen
- b) Oxygen, magnesium solution
- c) Magnesium,  $H_2O$
- d) Magnesium, oxygen

4. Choose incorrect option

- a) Change in reactant state
- b) Change in color
- c) Change in product state
- d) Change in temperature

### Section - B

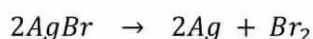
5. What is meant by a chemical reaction?

6. Name and state the law which is kept in mind when we balance chemical equations.

7. State one basic difference between a physical change and a chemical change.

8. On what basis is a chemical reaction balanced?

9. Write the essential condition for the following reaction to take place:



Write application of this reaction.

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## Section - C

10. Write the balanced chemical equations for the following chemical reactions:

- a. Hydrogen + Chlorine  $\rightarrow$  Hydrogen Chloride
- b. Lead + Copper Chloride  $\rightarrow$  Lead chloride + Copper
- c. Zinc oxide + Carbon  $\rightarrow$  Zinc + Carbon Monoxide

11. Write chemical equations for the reactions taking place when

- a. Iron reacts with steam
- b. Magnesium reacts with dil.  $HCl$
- c. Copper is heated in air

## Section - D

12.

- a. Define a balanced chemical equation. Why should an equation be balanced?
- b. Write a balanced chemical equation for the following reactions:
  - (i) Phosphorus burns in the presence of chlorine to form phosphorus pentachloride.
  - (ii) Burning of natural gas
  - (iii) The process of respiration.



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