

Science

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(Chapter – 1) (Chemical Reactions and Equations) (Practice Test 4)

(Class X)

Time: 1 Hour

Maximum Marks: 25

General Instructions:

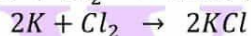
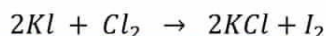
- This question paper contains four sections – A, B, C, and D. Each part is compulsory.
- Section – A has 4 MCQ of one mark each.
- Section – B has 5 questions of two marks each.
- Section – C has 2 questions of three marks each.
- Section – D has 1 question of five marks.

Section - A

- Choose the incorrect answer.
 - Reactants react and form a product which shows by an arrow placed between them
 - Products are written on the right side (RHS) and reactant (LHS)
 - Plus sign is always placed between two reactants
 - Arrow is pointed towards reactants
- Formula of magnesium oxide.
 - MnO_2
 - MgO
 - Mg_2O
 - MnO
- For the reaction,, are not indicated above and below the arrow in the equation.
 - Temperature
 - Pressure
 - Catalyst
 - Reactant
- Formula of slaked lime?
 - $Ca(OH)_2$
 - Ca_2CO_3
 - H_2O
 - CaO

Section - B

- Two reactions are given below:



Identify the type of reaction, giving justification in each case.

- It has been found that the marble of Taj is getting corroded due to development of industrial areas around it. Explain this fact by giving a chemical equation.
- What is a redox reaction? Identify the substances oxidized and the substance reduced in the following reactions:
$$MnO_2 + 4HCl \rightarrow MnCl_2 + Cl_2 + H_2O$$
$$CuO + H_2 \rightarrow Cu + H_2O$$
- Write balanced equation for the reaction between Mg and hydrochloric acid. Name the product obtained, identify the type of reaction.
- Using a balanced chemical equation explains the difference between a displacement reaction and a double displacement reaction.

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Section - C

10. State one example each characterized by the following along with a suitable chemical equation.
- Change in state,
 - Evolution of gas,
 - Change in temperature

- 11.
- Solid calcium oxide was taken in a container and water was added slowly to it.
 - Write the observations.
 - Write the chemical formula of the product formed.
 - What happens when carbon dioxide is bubbled through lime water (a) in small amounts (b) in excess?

Section - D

12. Write balanced chemical equation for the following statements:
- $NaOH$ solution is heated with zinc granules.
 - Excess of carbon dioxide is passed through lime water.
 - Dilute sulphuric acid is added to sodium carbonate.
 - Egg shell is dropped in hydrochloric acid.
 - Copper (II) oxide reacts with dilute hydrochloric acid.



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