

# Science

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## (Chapter - 1) (Chemical Reactions and Equations) (Practice Test 5) (Class X)

Time: 1 Hour

Maximum Marks: 25

### General Instructions:

- This question paper contains four sections – A, B, C, and D. Each part is compulsory.
- Section – A has 4 MCQ of one mark each.
- Section – B has 5 questions of two marks each.
- Section – C has 2 questions of three marks each.
- Section – D has 1 question out of five marks.

### Section - A

1. Use of slaked lime...
  - a) Used in medicines
  - b) Used in washing soda
  - c) Used in white washing walls
  - d) Used in pickle
2. Burning of natural gas is ..... Reaction
  - a) Endothermic
  - b) Decomposition
  - c) Photosynthesis
  - d) Exothermic
3. Carbohydrates broken down into .....
  - a) Water
  - b) Galactose
  - c) Fats
  - d) Glucose
4. Ferrous sulfate crystals have ..... in color.
  - a) Yellow
  - b) Blue
  - c) Green
  - d) Red

### Section - B

5. In electrolysis of water, why is the volume of gas collected over one electrode double that of gas collected over the other electrode.
6. If copper metal is heated over a flame it develops a coating. What is the color and composition of the coating?
7. On adding dilute hydrochloric acid to copper oxide powder, the solution formed is blue green. Predict the new compound formed which imparts a blue green color to the solution.
8. Why does the color of copper sulfate solution change when an iron pin is dipped in it?
9. What happens when an aqueous solution of sodium sulfate reacts with an aqueous solution of barium chloride? State the physical conditions of reactants in which the reaction between them will not take place. Write the balanced chemical equation for the reaction and name the type of reaction.

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(Class X)

### Section - C

10. (a) Why is it necessary to balance a chemical equation?  
(b) Write the balanced chemical equation for the following reactions:  
(i) Natural gas burns in air to form carbon dioxide and water.  
(ii) During respiration, glucose combines with oxygen and forms carbon dioxide and water along with the release of energy.
11. a. Can a combination reaction be an oxidation reaction?  
b. How will you test whether the gas evolved in a reaction is hydrogen?  
c. Why does copper not evolve hydrogen on reacting with dilute sulphuric acid?

### Section - D

12. Identify the type of reactions taking place in each of the following:
- Barium chloride solution is mixed with copper sulfate solution and white precipitate is formed.
  - On heating copper powder in China dish, the surface of copper powder turns black.
  - On heating green coloured ferrous sulfate crystals, reddish brown solid is left and smell of a gas having odor of burning sulfur is experienced.
  - Iron nails when left dipped in blue copper sulfate solution become reddish brown in color and the blue color of copper sulfate fades away.
  - Quick lime reacts vigorously with water releasing a large amount of heat.



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