

Science

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(Chapter - 2) (Acid Bases and Salts) (Practice Test 4)

(Class X)

Time: 60 minutes

M. M.: 25

General Instructions:

- This question paper contains four sections - A, B, C, and D. Each part is compulsory.
- Section - A has 4 MCQ of one mark each.
- Section - B has 5 questions of two mark each.
- Section - C has 2 questions of three mark each.
- Section - D has 1 question of five mark.
- There is no negative marking.

Section - A

1. All metal carbonates and hydrogen carbonates reacts with acids to give
(A) Salt, Carbon Dioxide, Water
(B) Salt, water, acid
(C) Salt water, base
(D) Salt water, hydrogen gas
2. Copper (II) chloride color
(A) Green yellow
(B) Green purple
(C) Blue green
(D) Reddish pink
3. Metallic oxides are
(A) Salt
(B) In powder formed
(C) Liquid in nature
(D) Basic oxides
4. Non-metallic oxides are in nature
(A) Basic
(B) Acidic
(C) Salty
(D) Solid

Section - B

5. Write a balanced chemical equation for the reaction between sodium carbonate and hydrochloric acid indicating the physical state of reactants and the products.
6. Oxides of metals are basic while those of non-metals are acidic. Explain.
7. What is the difference between slaked lime and lime water?
8. Write a balanced chemical equation for the neutralization reaction, mentioning the physical state of reactants and products.
9. Give suitable reasons to justify the following statement:
An aqueous solution of sodium chloride is neutral but an aqueous solution of sodium metal is basic.

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Section - C

10. Explain the action of dilute hydrochloric acid on the following with suitable chemical equations:

(A) Magnesium ribbon, (B) Sodium hydroxide, (C) Crushed egg shells.

11. You are provided with magnesium ribbon and sulfur powder. Explain with the help of activity that metal oxides are basic and oxides of non-metals are acidic in nature.

Section - D

12. Write chemical equations when zinc granules react with

- (A) Sulphuric acid,
- (B) Hydrochloric acid,
- (C) Aluminum chloride,
- (D) Sodium hydroxide,
- (E) Nitric acid



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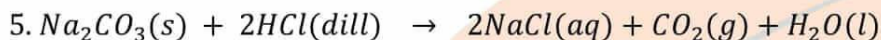
(Class X)

Hints and Answers

Section - A

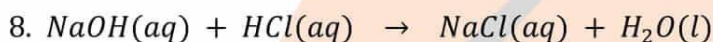
1. (a) Salt, Carbon Dioxide, water.
2. (c) Blue green.
3. (d) Basic oxides.
4. (b) Acidic.

Section - B

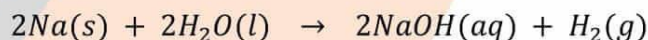


6. Metal oxides dissolve in water to form base basic in nature. On the other hand non-metals dissolve in water to form acids, acidic in nature.

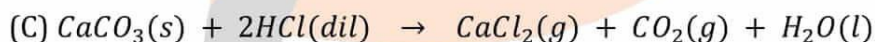
7. The solid $Ca(OH)_2$ is slaked lime whereas clear solution of $Ca(OH)_2$ in water is lime water.



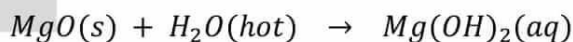
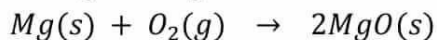
9. Sodium chloride is made up of a strong base, $NaOH$ and a strong acid, HCl . Therefore, its aqueous solution is neutral in nature. Sodium metal reacts with water to form $NaOH$ (Base) and H_2 gas:



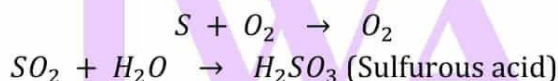
10.



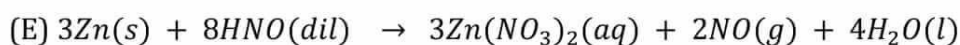
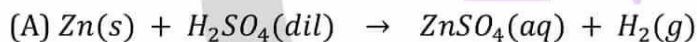
11. Burn magnesium ribbon with the help of tongs to form white ash. Dissolve the ash in hot water. Add red litmus which turns blue, showing that MgO is a basic oxide.



Heat sulfur taken in an iron spatula and pass the gas through water. Add blue litmus into it. It will turn red showing SO_2 is an acidic oxide.



12.



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