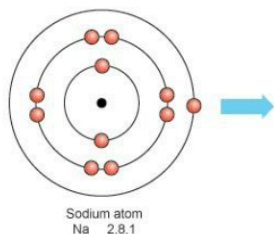


**Question 1:**

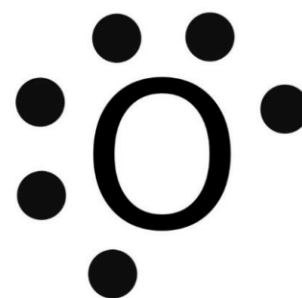
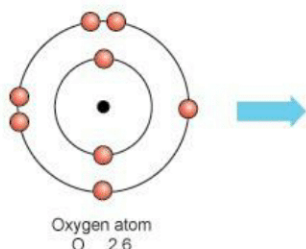
- (i) Write the electron-dot structures for sodium, oxygen and magnesium.
- (ii) Show the formation of Na<sub>2</sub>O and MgO by the transfer of electrons.
- (iii) What are the ions present in these compounds?

**Answer 1:**

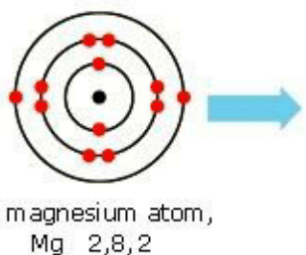
**(i) Electron - dot structure for Sodium:**



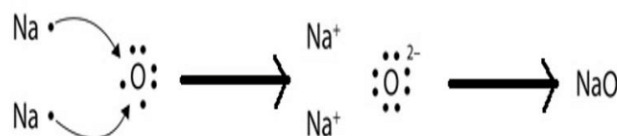
**Electron - dot structure for Oxygen:**



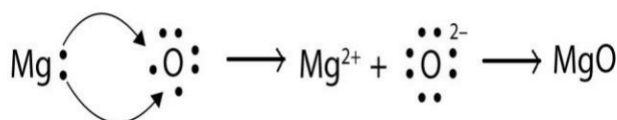
**Electron - dot structure for Magnesium:**



**(ii). Formation of Na<sub>2</sub>O by transfer of electron:**



**Formation of MgO by transfer of electron:**



**(iii). Ions present in these compounds are Mg<sup>2+</sup>, O<sup>2-</sup> and Na<sup>+</sup>.**

**Question 2:**

Why do ionic compounds have high melting points?

**Answer 2:**

Ionic compounds have high melting and boiling points. Because ionic compounds are formed by the attraction force of two opposite ions and a considerable amount of energy is required to break this strong inter-ionic attraction.